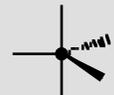
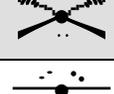
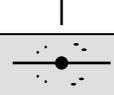


Shaded squares represent geometries which give non-polar molecules when all substituents X are identical.

Total # of Groups of e^-	Electron Pair Geometry (Hybridization)	Approximate Bond Angle	# of Bonding Directions (# of X)	# of Lone Pairs (# of E)	Geometry Name (VSEPR class)	Shape	Examples
2	linear (sp)	180°	2	0	linear (AX_2)		BeH ₂ , CO ₂
3	trigonal planar (sp^2)	120°	3	0	trigonal planar (AX_3)		BF ₃ , NO ₃ ⁻
			2	1	bent (AX_2E)		SO ₂
4	tetrahedral (sp^3)	109.5°	4	0	tetrahedral (AX_4)		CH ₄
			3	1	trigonal pyramidal (AX_3E)		NH ₃
			2	2	bent (AX_2E_2)		H ₂ O
5	trigonal bipyramidal (sp^3d)	120° (in plane) & 90° (above & below)	5	0	trigonal bipyramidal (AX_5)		PCl ₅
			4	1	seesaw (AX_4E)		SF ₄
			3	2	T-shaped (AX_3E_2)		ClF ₃
			2	3	linear (AX_2E_3)		XeF ₂
6	octahedral (sp^3d^2)	90°	6	0	octahedral (AX_6)		SF ₆
			5	1	square pyramidal (AX_5E)		BrF ₅
			4	2	square planar (AX_4E_2)		XeF ₄
			3	3	T-shaped (AX_3E_3)		
			2	4	linear (AX_2E_4)		